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 $24.3 + 12 + 3(16) = 84.3 \text{ g/mol MgCO}_3$
 $+ 2 \text{ H}^+ = \text{Mg}^{2+} + \text{H}_2\text{CO}_3$ EW = $84.3 / 2$
 $= 42.15 \text{ g/equiv}$ (b) NaNO_3 FW = $23 +$
 $14 + 3(16) = 85 \text{ g/mol NaNO}_3 + \text{H}^+ =$
 $\text{Na}^+ + \text{HNO}_3$ EW = $85 / 1 = 85 \text{ g/equiv}$

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$$(c) \text{CO}_2 \text{ FW} = 12 + 2(16) = 44 \text{ g/mol}$$

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